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# Room temperature one-step synthesis of microarrays of N-doped flower-like anatase TiO<sub>2</sub> composed of well-defined multilayer nanoflakes by Ti anodization

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## Abstract

Microarrays of N-doped flower-like TiO<sub>2</sub> composed of well-defined multilayer nanoflakes were synthesized at room temperature by electrochemical anodization of Ti in NH<sub>4</sub>F aqueous solution. The TiO<sub>2</sub> flowers were of good anatase crystallinity. The effects of anodizing time, applied voltage and NH<sub>4</sub>F concentration on the flower-like morphology were systematically examined. It was found that the morphologies of the anodized Ti were related to the anodizing time and NH<sub>4</sub>F concentration. The size and density of the TiO<sub>2</sub> flowers could be tuned by changing the applied voltage. The obtained N-doped flower-like TiO<sub>2</sub> microarrays exhibited intense absorption in wavelengths ranging from 320 to 800 nm. Under both UV and visible light irradiation, the photocatalytic activity of the N-doped flower-like TiO<sub>2</sub> microarrays in the oxidation of methyl orange showed a significant increase compared with that of commercial P25 TiO<sub>2</sub> film.

## 1. Introduction

TiO<sub>2</sub> photocatalysts have attracted considerable attention as highly functional materials due to their semiconductor nature, long term stability, and low-cost preparation for a variety of applications, including the purification of air and water [1, 2], splitting of water into hydrogen [3], and dye-sensitized solar cells [4, 5]. TiO<sub>2</sub> with different morphologies possesses various properties and applications due to its specific size and shape [6]. To this end, the controlled synthesis of TiO<sub>2</sub> with different sizes and shapes has become an increasingly significant topic for material science and technology. Recent research on the synthesis of micro- and nano-materials has shown that these kinds of TiO<sub>2</sub> architectures may promise potential use in catalysts and optoelectronics [7–9].

Anatase TiO<sub>2</sub> exhibits a higher activity than its rutile counterpart because its conduction band is at a higher

energy [1]. Various approaches such as hydrothermal treatment [6, 10], the sol–gel method [11, 12] and calcination processes [13, 14] have been developed to synthesize anatase TiO<sub>2</sub> structures at low temperature, but most of them unavoidably involve the use of either templates or complicated procedures. When anatase TiO<sub>2</sub> is prepared by hydrothermal treatment or sol–gel synthesis, the crystallization usually results in highly reduced properties, structural damage or particle aggregation [15–17]. Similarly, anatase TiO<sub>2</sub> obtained by calcination generally produces large-sized particles, which leads to structural collapse [15]. For a long time, suspensions or compacted nanoparticulate films have mainly been used, based mainly on the commercial Degussa P25 TiO<sub>2</sub> nanopowder.

It is well known that the industrial application of the suspended system in the form of colloidal and particulate TiO<sub>2</sub> in photodegradation of pollutants is seriously limited due

to the difficulty in separating TiO<sub>2</sub> from water or pollutants and recycling the photocatalyst, the particle aggregation especially at high concentration, and the problematic use in continuous flow systems. To overcome these technical problems, various methods have been developed to immobilize TiO<sub>2</sub> films on solid supportive substrates. However, the efficiency of the immobilized systems is much lower than that of the corresponding slurries, because of the inevitable reduction of the overall surface active area associated with the catalyst immobilization. In the past decade, electrochemical anodization has been demonstrated to be a convenient technique for fabricating a variety of TiO<sub>2</sub> films on Ti substrates [18–24]. By tuning the electrochemical parameters, such as the composition and concentration of electrolytes, anodizing voltage, and anodizing time, TiO<sub>2</sub> films with different shapes and sizes can be fabricated [21–24]. However, the as-fabricated TiO<sub>2</sub> films are amorphous [25]. To convert amorphous TiO<sub>2</sub> into anatase phase, thermal annealing at high temperature is necessary [21–24, 26, 27]. Very recently, a few modified approaches were demonstrated to grow partially crystalline, anatase TiO<sub>2</sub> nanotubes at room temperature [28], and totally anatase TiO<sub>2</sub> nanotubes at 60 °C by anodizing Ti substrates [29]. In particular, Grimes *et al* reported a room temperature one-step polyol synthesis of anatase TiO<sub>2</sub> nanotube arrays by an optimized anodization process [30]. In this paper, we report the room temperature one-step synthesis of microarrays of N-doped flower-like anatase TiO<sub>2</sub> composed of well-defined multilayer nanoflakes by a simple electrochemical anodization method. To the best of our knowledge, this is the first study of room temperature synthesis of anatase TiO<sub>2</sub> in an aqueous electrolyte. Notably, this is also the first report on one-step synthesis of N-doped flower-like TiO<sub>2</sub> by Ti anodization. The effects of anodizing time, applied voltage and NH<sub>4</sub>F concentration on the morphology of the TiO<sub>2</sub> structures were examined in detail. The light absorption property and photocatalytic activity of the N-doped flower-like anatase TiO<sub>2</sub> microarrays were investigated. The ability to synthesize N-doped anatase TiO<sub>2</sub> materials at room temperature makes them very attractive for practical applications in environmental purification, water photoelectrolysis and solar energy conversion.

## 2. Experimental details

Titanium foils of 0.1 mm in thickness and a purity of 99.6% were used. Prior to anodization, the Ti foils (1 cm × 1 cm) were ultrasonically cleaned with anhydrous ethanol and distilled water successively, followed by rinsing with DI water and drying in air. The anodization was carried out in a two-electrode electrochemical cell with titanium foil as the working electrode and platinum foil as the counter electrode under constant voltage applied by a direct current (DC) power supply (HY1712-5S, Huaian Yaguang Electronics Co., Ltd). The distance between the working electrode and the counter electrode was 2.5 cm and the electrolyte was an NH<sub>4</sub>F aqueous solution. All experiments were carried out at room temperature (~25 °C). After anodization, the samples were immediately rinsed with deionized water, and then dried at

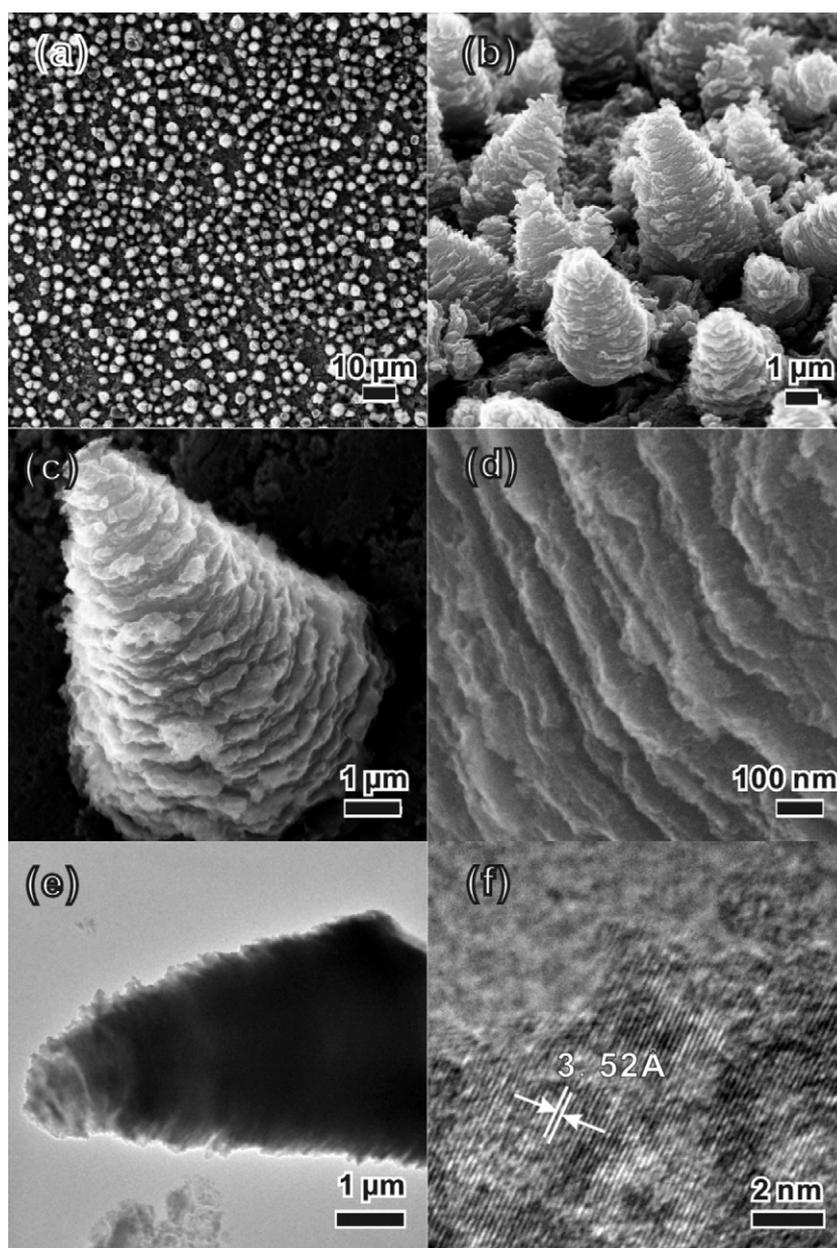
ambient temperature. For comparison, a commercial Degussa P25 TiO<sub>2</sub> powder (a mixture of anatase and rutile) was used to prepare thin particulate film on the Ti foil. P25 film was prepared by dispersing into deionized water first and then depositing on the Ti foil, followed by a heat treatment at 200 °C for 30 min [31]. This procedure was repeated several times until the thickness of the P25 film reached ca. 10 μm.

The structure and morphology of the anodized samples were examined using an LEO-1530 field-emission scanning electron microscope (FESEM) and a Tecnai-F30 high-resolution transmission electron microscope (HRTEM). The crystalline phases of the samples were detected by a Philips X'pert x-ray diffractometer (XRD) with Cu Kα radiation. The composition of the sample was identified by x-ray photoelectron spectroscopy (XPS; VG, Physical Electronics Quantum 2000 Scanning ESCA Microprobe, Al Kα radiation). UV–visible diffuse reflectance spectra of the samples were recorded using a UV–vis–NIR spectrophotometer (Varian Cary 5000).

The photocatalytic activity of the samples was evaluated by the degradation of the methyl orange (MO) azo dye. The photocatalytic experiments were carried out in a quartz glass reactor. The reactor was equipped with a water jacket to control the reaction temperature. A 200 W high-pressure mercury lamp emitting at a wavelength of 365 nm was employed as the UV light source or a 500 W tungsten–halogen lamp ( $\lambda \geq 520$  nm) was used to produce the simulated sunlight. All experiments were performed under continuous stirring, using 30 ml of 20 mg l<sup>-1</sup> MO aqueous solution, and air was bubbled through the gas disperser into the reactor. The absorbance changes of the MO were monitored using a UV–vis spectrophotometer (Unico UV-2102 PC) at a wavelength of 507 nm.

## 3. Results and discussion

A representative large-area FESEM image of the TiO<sub>2</sub> microarray synthesized in 0.01 M NH<sub>4</sub>F electrolyte at 50 V for 2 h is shown in figure 1(a). A closer examination reveals that vertically oriented bamboo-shoot-like TiO<sub>2</sub> flowers about 2.5–7.0 μm in diameter and 2.1–10.3 μm in height grew on the Ti substrate (figure 1(b)), and the average diameter of the flower bottom and height of the flower are approximately 4.5 μm and 5.2 μm, respectively. Figure 1(c) shows the FESEM image of an individual flower. Further scrutiny of the flowers reveals that the flower-like TiO<sub>2</sub> consists of well-defined layered nanoflakes (figure 1(d); close-up of nanoflower shown in figure 1(c)). The corresponding TEM image shows that the flower-like TiO<sub>2</sub> has a hierarchical structure (figure 1(e)). An HRTEM image of the nanoflake reveals a lattice spacing of 0.352 nm, corresponding to the (101) plane of anatase TiO<sub>2</sub> (figure 1(f)) [32]. The corresponding x-ray diffraction (XRD) pattern is shown in figure 2. The diffraction peaks can be fully indexed to anatase phase TiO<sub>2</sub> (JCPDS 21-1272), indicating that the as-synthesized product was pure anatase TiO<sub>2</sub>. The size of TiO<sub>2</sub> crystallites was estimated to be 12 nm from the full-width at half-maximum (FWHM) of the (101) diffraction peak using the Scherrer equation. Some works



**Figure 1.** FESEM images of the as-synthesized  $\text{TiO}_2$ : (a) a top view, (b) a 3D view, (c) an individual  $\text{TiO}_2$  flower, (d) close-up of (c), (e) TEM image and (f) HRTEM image from an individual nanoflake.

have reported the formation of  $\text{TiO}_2$  nanotube arrays by the anodization of Ti in various electrolytes containing  $\text{NH}_4\text{F}$ , including 1 M  $(\text{NH}_4)_2\text{SO}_4$  [33, 34], 1 M  $(\text{NH}_4)_2\text{H}_2\text{PO}_4$  [35], glycerol electrolyte [36] and ‘water-free’  $\text{CH}_3\text{COOH}$  [37], which respectively contain 0.5 wt%  $\text{NH}_4\text{F}$ . In contrast, a simple aqueous solution containing 0.01 M  $\text{NH}_4\text{F}$  was used in our work; therefore, only flower-like  $\text{TiO}_2$  arrays, but not  $\text{TiO}_2$  nanotube arrays, were observed in the experiments.

In order to elucidate the formation process of the flower-like  $\text{TiO}_2$  microarrays, we investigated the influence of the reaction time when the Ti foils were anodized in the  $\text{NH}_4\text{F}$  solution at a constant anodization potential of 50 V. The FESEM images of the samples yielded at different time periods are shown in figure 3. After 10 min anodization (figure 3(a)),

the surface of the Ti substrate was covered by an initial porous oxide layer [18, 21]. As time increased to 30 min (figure 3(b)), small protrusions appeared randomly on the Ti surface and were considered as the buds of flower-like  $\text{TiO}_2$ . As the anodization time progressed to 1 h (figure 3(c)), the protrusions were converted into bamboo-shoot-like  $\text{TiO}_2$  (referred to as flower-like  $\text{TiO}_2$ ), and the number of formed flower-like  $\text{TiO}_2$  increased. Further increase of the anodization time to 2 h (figure 3(d)) led to increased coverage of the flower-like  $\text{TiO}_2$  on the Ti substrate. Finally, the self-organized and vertically oriented flower-like  $\text{TiO}_2$  microarray was thus obtained.

Figure 4 shows the current–time response behavior of Ti anodized at 50 V in 0.01 M  $\text{NH}_4\text{F}$  aqueous solution. Such a curve shape, showing a decrease/increase/decrease sequence,

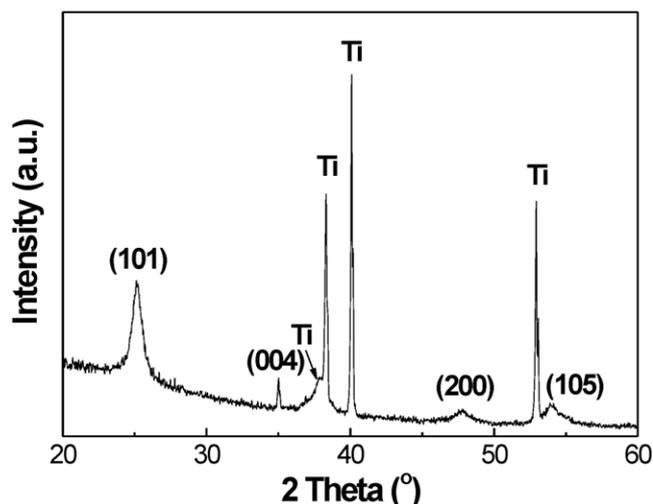


Figure 2. XRD patterns of the as-synthesized TiO<sub>2</sub> microarray.

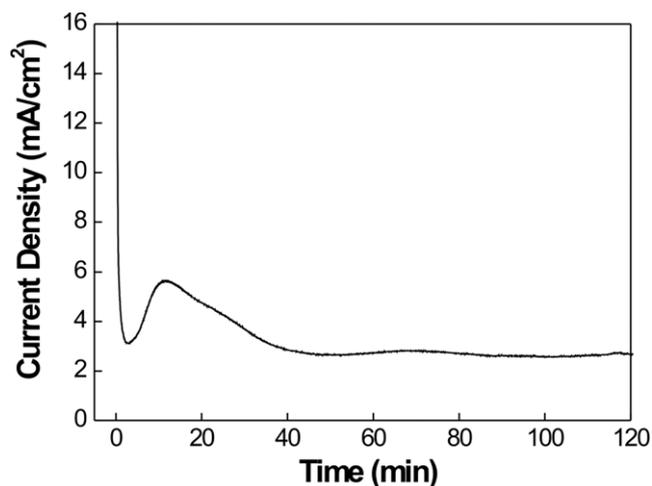


Figure 4. Current density versus time during the anodization of Ti foil in a 0.01 M NH<sub>4</sub>F electrolyte at 50 V.

has already been observed for self-organized formation of TiO<sub>2</sub> nanotubes [38, 28]. Initially, the current density dropped sharply due to the formation of a thin oxide layer [39, 40], and then increased to a maximum due to the pitting of the oxide layer which resulted from the field-assisted dissolution [41]. Some quite small TiO<sub>2</sub> grains with defects formed in the preferential sites (marked with arrows in figure 3(a)). Then the current density started to decrease again due to an increase

in the porous depth along with the local dissolution of Ti. The TiO<sub>2</sub> grains tended to coalesce due to the high activity of their respective surface defects [42, 43]. The crystalline transformation temperature for smaller TiO<sub>2</sub> particles is relatively lower [44]. As a consequence, the crystalline TiO<sub>2</sub> nanoflakes formed in the suitable local oxidation environment. In the progress of TiO<sub>2</sub> nanoflake corrosion and precipitation on the surface, crystalline TiO<sub>2</sub> nanoflakes continuously

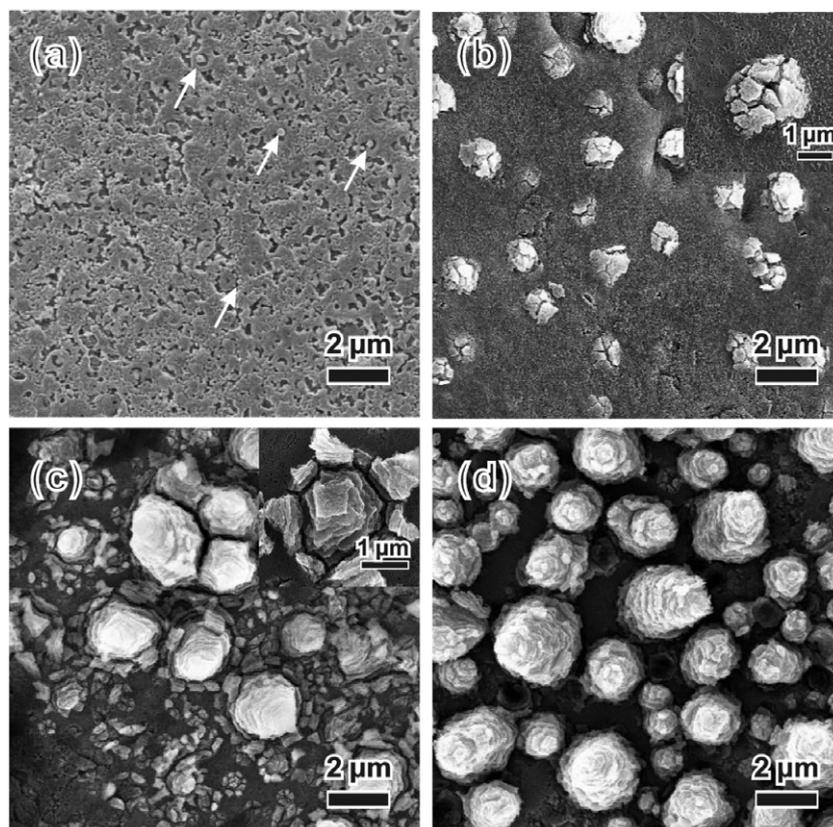
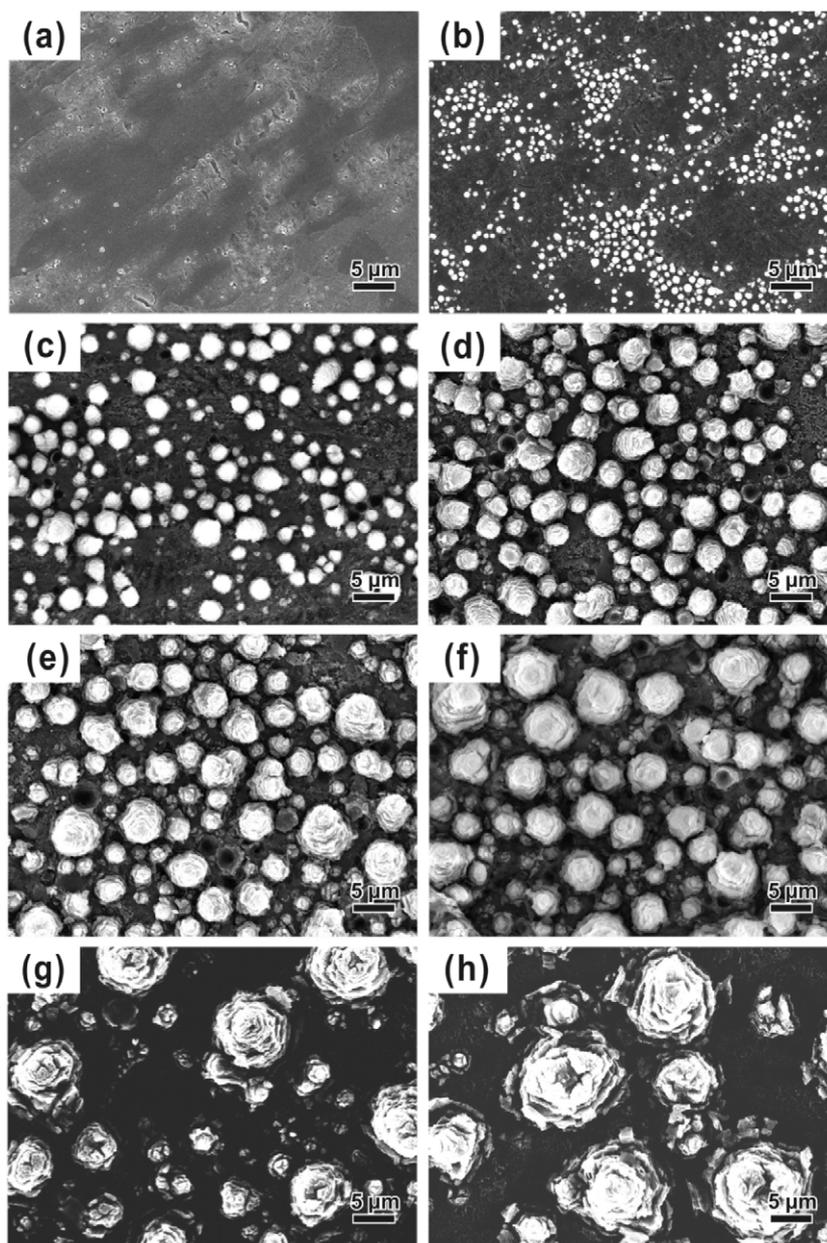


Figure 3. FESEM images showing the evolution of the flower-like structured TiO<sub>2</sub> when the Ti foil was anodized in an NH<sub>4</sub>F solution at a constant potential of 50 V for 10 min (a), 30 min (b), 1 h (c) and 2 h (d).

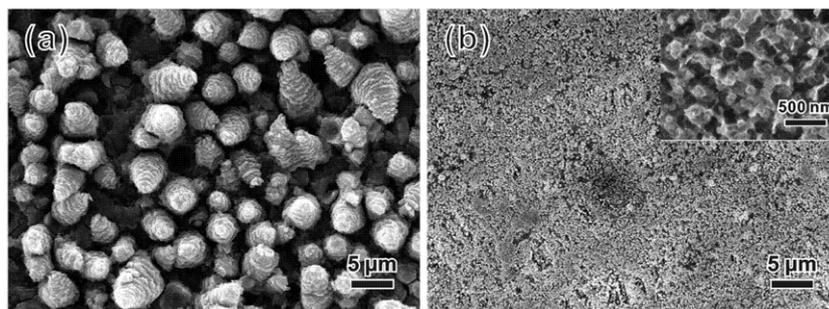


**Figure 5.** FESEM images of flower-like anatase TiO<sub>2</sub> anodized at different voltages: (a) 20 V, (b) 30 V, (c) 40 V, (d) 50 V, (e) 60 V, (f) 70 V, (g) 80 V, and (h) 90 V.

formed at the bottom of the pores and gradually layered up in the direction normal to the Ti substrate. Some bud-like TiO<sub>2</sub> were shaped by aggregating nanoflakes, which is clearly evident in figure 3(b). As the anodization progressed, flower-like TiO<sub>2</sub> appeared and their size gradually increased with the prolonged anodization time (figure 3(c)). Finally, due to the dynamic balance between the field-assisted oxidation and dissolution as well as the chemical dissolution, the growth of the flower-like TiO<sub>2</sub> was limited to a certain shape and size, and the flower-like anatase TiO<sub>2</sub> with multilayered nanoflakes were thus yielded (figure 3(d)).

In order to assess the influence of other key parameters on the surface morphology, a set of experiments was performed by systematically varying the applied voltages and the NH<sub>4</sub>F

concentration. Figures 5(a)–(h) show FESEM images of the flower-like structures obtained at different anodizing voltages (i.e., from 20 to 90 V with a 10 V increment) in 0.01 M NH<sub>4</sub>F solution for 2 h. At low anodizing voltage (20 V; figure 5(a)) only a few flower-like particles formed. As the voltage increased from 30 to 50 V (figures 5(b)–(d)), the amount of flower-like particles and their size began to increase. When the applied voltage was at 60 and 70 V (figures 5(e) and (f)), no obvious changes in the size and density of flower-like particles could be observed. As the voltage further increased to above 70 V (figures 5(g) and (h) at 80 V and 90 V, respectively), the size of the flower-like particles continued to increase, while the density decreased. We propose that the formation of the flower-like structures may be related to the low concentration



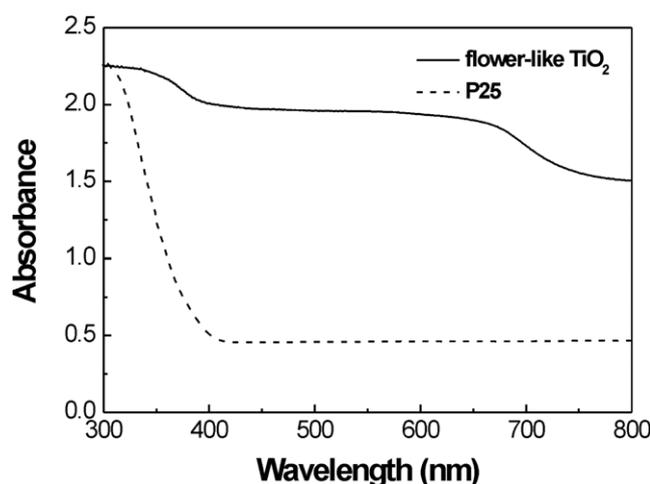
**Figure 6.** FESEM images of flower-like anatase TiO<sub>2</sub> formed at 50 V for 2 h in 0.05 M NH<sub>4</sub>F (a) and 1.0 M NH<sub>4</sub>F (b).

of aqueous electrolyte. When the voltage was very low, nearly dense TiO<sub>2</sub> film was obtained which was due to the low etching rate of TiO<sub>2</sub> (mainly electric field induced dissolution); as the voltage increased, flower-like structures started to appear and the structure density decreased with the increase in voltage. As a result, flower-like TiO<sub>2</sub> can be obtained in the range from 30 to 90 V.

Figures 6(a) and (b) represent the FESEM images of the samples formed at 50 V for 2 h in 0.05 M and 1.0 M NH<sub>4</sub>F solution, respectively. In the 0.05 M NH<sub>4</sub>F solution self-organized flower-like TiO<sub>2</sub> can still be achieved. Anodization in the 1.0 M NH<sub>4</sub>F solution, however, resulted in a porous structure. These observations confirmed that there was a chemical equilibrium between the formation and dissolution of TiO<sub>2</sub> during the formation of flower-like TiO<sub>2</sub>. When anodization was carried out in low fluorine ion (F<sup>-</sup>) concentration (figure 6(a)), the formation rate of flower-like TiO<sub>2</sub> was faster than the rate of its dissolution, thereby leading to the formation of flower-like morphology. By contrast, when anodization was performed at high F<sup>-</sup> concentration (figure 6(b)), the dissolution of TiO<sub>2</sub> to form pores occurred according to the reaction [33]



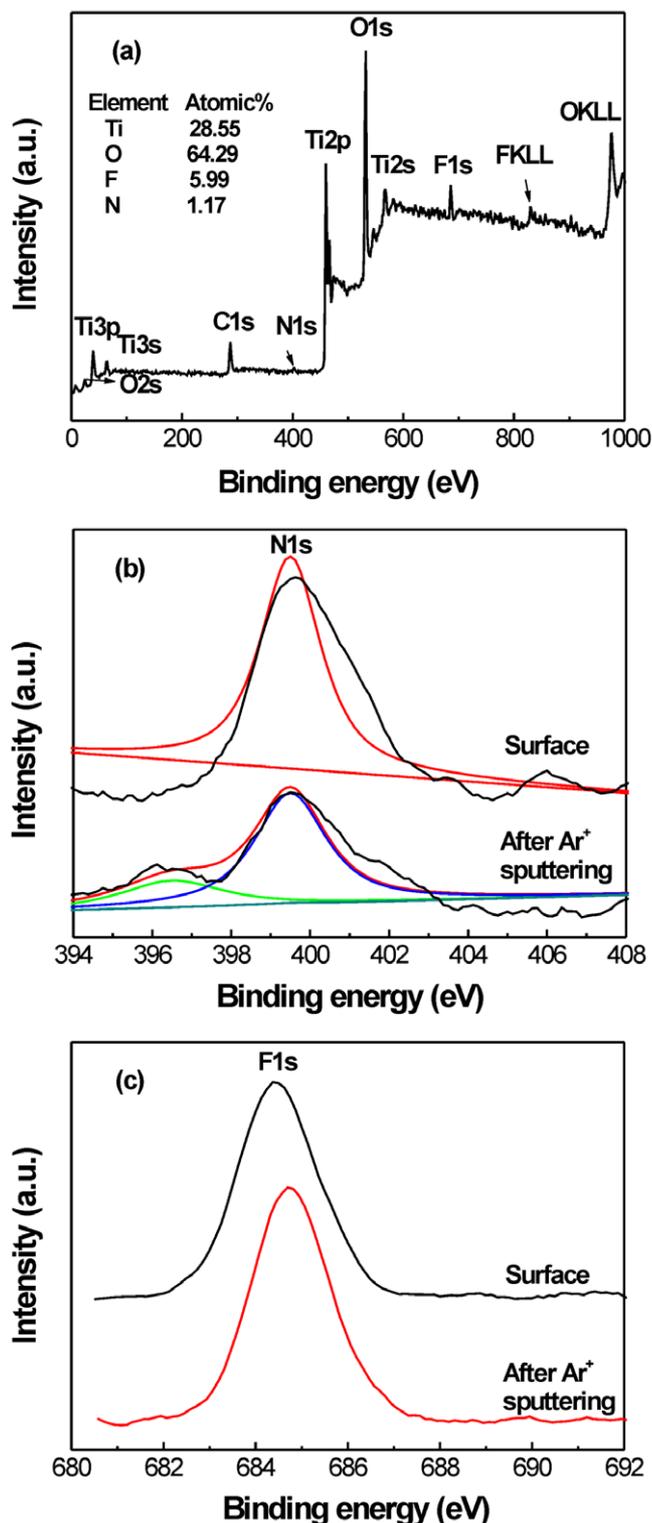
We further studied the light absorption property of the flower-like anatase TiO<sub>2</sub> microarray. For comparison, P25, one of the best commercial TiO<sub>2</sub> photocatalysts, was chosen as the benchmark. Figure 7 shows a comparison of UV-vis diffuse reflectance spectra for the flower-like anatase TiO<sub>2</sub> microarray anodized in 0.01 M NH<sub>4</sub>F electrolyte at 50 V for 2 h in which the TiO<sub>2</sub> flowers show relatively uniform size and larger density and a P25 film whose thickness was the same as the largest height of the TiO<sub>2</sub> flowers (approximately 10 μm). A significant absorption of P25 film at wavelength shorter than 400 nm can be assigned to the intrinsic band gap absorption of anatase phase TiO<sub>2</sub> (≈3.2 eV). By contrast, the flower-like anatase TiO<sub>2</sub> microarray exhibited a much stronger absorption in the range from 320 to 800 nm and a remarkable red shift in the band gap transition. A previous report showed that TiO<sub>2</sub> nanotube arrays prepared by the anodization of Ti in an electrolyte without NH<sub>4</sub>F, such as HF aqueous solution, did not exhibit visible light absorption [45]. However, in our experiment, the flower-like anatase TiO<sub>2</sub> microarray prepared by the anodization of Ti in NH<sub>4</sub>F aqueous solution without any



**Figure 7.** UV-vis diffuse reflectance spectra of the N-doped flower-like anatase TiO<sub>2</sub> and the P25 film.

organic compound exhibited stronger visible light absorption, indicating that the presence of NH<sub>4</sub><sup>+</sup> in the electrolyte plays a key role in the visible light absorption. The intense visible light response of the flower-like anatase TiO<sub>2</sub> microarray could be primarily attributed to the doping of N in the TiO<sub>2</sub>. In addition, the unique flower-like morphology of anatase TiO<sub>2</sub> and its hierarchical structures composed of layered nanoflakes could lead to more reflectance of the light.

The form of the N in the flower-like anatase TiO<sub>2</sub> was determined by XPS. The XPS spectrum (figure 8(a)) indicated that the as-prepared flower-like anatase TiO<sub>2</sub> contained Ti, O, N, F and a trace amount of C. The atomic ratio of Ti:O was 1:2.25, which is much larger than stoichiometric TiO<sub>2</sub>. To identify the chemical state of O, Ar<sup>+</sup> sputtering was performed to remove a 5 nm surface layer. After Ar<sup>+</sup> sputtering, the atomic ratio of Ti:O was 1:2.0, indicating that the excess oxygen at the surface of the flower-like TiO<sub>2</sub> should be attributed to surface adsorbed oxygen. The existence of the C element can be ascribed to the adventitious hydrocarbon from the XPS instrument itself. In order to uncover the origin of the N and F signals, high-resolution XPS spectra of the N 1s region and the F 1s region of the sample anodized in 0.01 M NH<sub>4</sub>F electrolyte at 50 V for 2 h were performed and are shown in figures 8(b) and (c), respectively. A broad peak in the range from 396 to 404 eV (figure 8(b)), with a maximum



**Figure 8.** XPS spectra of (a) the flower-like TiO<sub>2</sub> over a large energy range, (b) N 1s and (c) F 1s.

(This figure is in colour only in the electronic version)

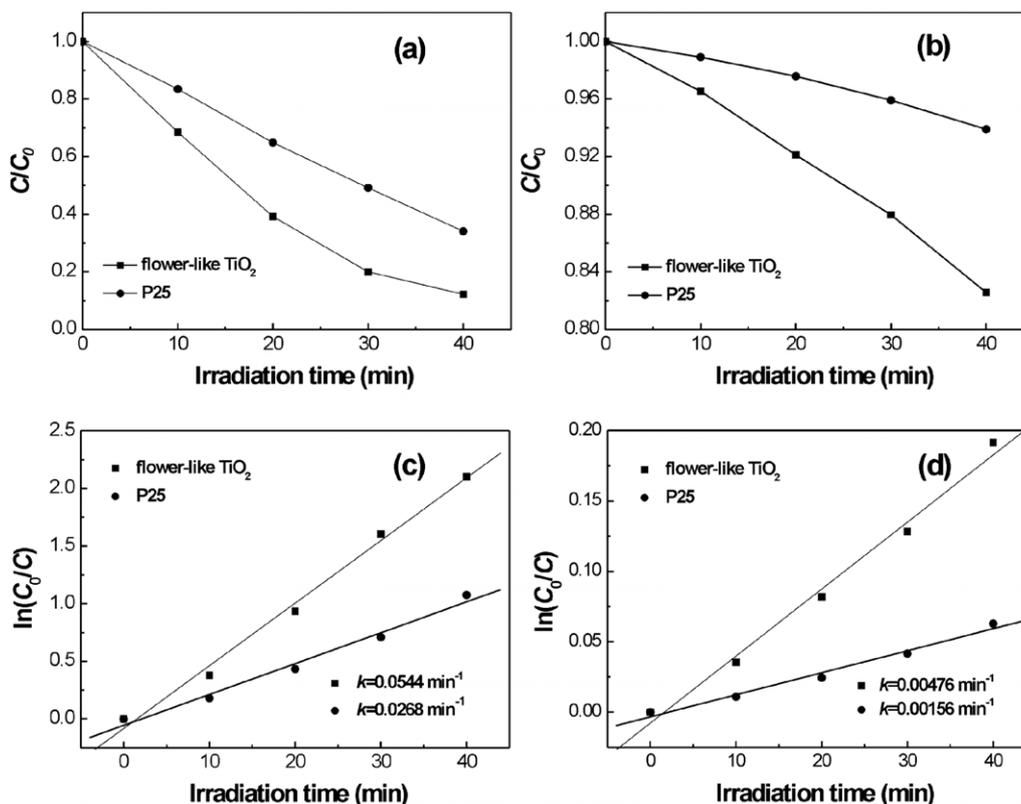
located at the binding energy of 399.6 eV, was clearly evident, which can be attributed to N 1s [46, 47]. A similar peak at nearly 399.6 eV was also identified in other literature [48, 49] from the ammonia-treated TiO<sub>2</sub> and could be assigned to the

N 1s band of interstitial doped N atoms [50, 51]. After Ar<sup>+</sup> sputtering, this peak became a little weak. It is noteworthy that a weak peak at 396–397 eV was detected and attributed to the N 1s band of substitutional nitrogen atoms that can be identified with the atomic β-N states [43]. Therefore, N from the NH<sub>4</sub>F solution was doped into the TiO<sub>2</sub> crystal lattice. As shown in figure 8(c), the peak located at 684.4 eV can be attributed to the F<sup>-</sup> ions physically adsorbed on the surface of the TiO<sub>2</sub> [8, 34, 52]. After Ar<sup>+</sup> sputtering, this peak became a little weak and no new peak appeared, indicating that the F<sup>-</sup> ions in the electrolyte were only physically adsorbed on the surface of TiO<sub>2</sub>. Accordingly, the significant visible light absorption of the flower-like anatase TiO<sub>2</sub> mainly resulted from the doping of N in TiO<sub>2</sub>. In this context, it may be expected that the as-synthesized flower-like anatase TiO<sub>2</sub> microarrays may display an outstanding visible light photocatalytic activity.

To this end, the photocatalytic activity of the N-doped flower-like anatase TiO<sub>2</sub> microarray was examined using methyl orange (MO) as a probe molecule. We tested the photocatalytic activity of four kinds of TiO<sub>2</sub> microarrays anodized in 0.01 M NH<sub>4</sub>F aqueous solution at 40, 50, 60, and 70 V for 2 h. It was found that the sample anodized in 0.01 M NH<sub>4</sub>F electrolyte at 50 V for 2 h exhibited the highest photocatalytic activity. Figures 9(a) and (b) show the photodegradation curves of MO in the presence of the N-doped flower-like anatase TiO<sub>2</sub> microarray which exhibited the highest photocatalytic activity and P25 film under UV and visible light irradiation, respectively. The linear relationship of  $\ln(c/c_0)$  versus time (figures 9(c) and (d)) shows that the photocatalytic degradation of MO followed a pseudo-first-order expression [53]:

$$\ln(C_0/C) = kt \quad (2)$$

where  $k$  is the apparent first-order rate constant ( $\text{min}^{-1}$ ),  $C_0$  and  $C$  are the initial and reaction concentrations of MO dye, respectively. Under UV irradiation (figure 9(c)), the apparent rate constant for the N–F-doped flower-like anatase TiO<sub>2</sub> was  $0.0544 \text{ min}^{-1}$ , which was about two times larger than that of the P25 film,  $0.0268 \text{ min}^{-1}$ . The unique morphology and hierarchical structure of flower-like anatase TiO<sub>2</sub> may play an important role in enhancing the UV photocatalytic activity of TiO<sub>2</sub>. The hierarchical structure composed of nanoflakes allowed an effective light-scattering inside the gaps between nanoflakes, resulting in the photogenerated electrons and holes to participating effectively in the photocatalytic degradation of contaminants [46]. Moreover, the micro–nano architecture of flower-like anatase TiO<sub>2</sub> shortened the path of photogenerated charge transfer at the interface [54]. Under visible light irradiation (figure 9(d)), the apparent rate constant for the N-doped flower-like anatase TiO<sub>2</sub> was  $0.00476 \text{ min}^{-1}$ , which was more than three times larger than that of the P25 film,  $0.00156 \text{ min}^{-1}$ , further confirming that the flower-like anatase TiO<sub>2</sub> exhibited excellent visible light photocatalytic activity for promising environmental applications. To check the potential reusability and stability of the N–F-doped flower-like anatase TiO<sub>2</sub> photocatalyst in the reaction medium, we performed a series of five photocatalytic cycles using the N-doped flower-like anatase TiO<sub>2</sub> microarray anodized in 0.01 M NH<sub>4</sub>F



**Figure 9.** Photodegradation curves of MO assisted by the N-doped flower-like anatase TiO<sub>2</sub> and the P25 film under UV (a) and visible light (b) irradiation. The fitting results using the pseudo-first-order reaction under UV and visible light irradiation are shown in (c) and (d), respectively.

electrolyte at 50 V for 2 h. The MO photodegradation as well as the first-order reaction constant  $k$  did not vary significantly, and the results were reproducible within 5%, indicating that the photocatalyst can be safely reused. In the present study, the N-doped flower-like anatase TiO<sub>2</sub> microarray was synthesized by anodizing Ti foil in NH<sub>4</sub>F aqueous solution at room temperature and exhibited excellent visible light photocatalytic activity. As a result, no thermal annealing is necessary and the resulting N-F-doped flower-like anatase TiO<sub>2</sub> microarrays can be readily used for practical applications.

#### 4. Conclusion

In summary, we reported the room temperature one-step synthesis of microarrays of N-doped flower-like anatase TiO<sub>2</sub> composed of well-defined multilayer nanoflakes by a simple, low-cost electrochemical anodization method. The applied voltage and the concentration of NH<sub>4</sub>F electrolyte played a crucial role in the growth of the flower-like anatase TiO<sub>2</sub> microarrays on the Ti substrate. The resulting N-doped flower-like anatase TiO<sub>2</sub> microarrays exhibited strong visible light response and excellent photocatalytic activity in the degradation of organic pollutants. The N-doped flower-like anatase TiO<sub>2</sub> material synthesized in the present study was directly grown on the Ti substrate at room temperature, dispensing with high temperature thermal annealing, thereby giving it great potential in environmental purification, water photoelectrolysis, and solar energy utilization.

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